

# Geothermometry in fluid phase in geothermal exploration

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## **Target(s) of exploration**

Geological and geochemical parameters, in terms of qualitative observations and quantitative estimates (geothermometry s.s.) should drive, together, to the formation of a conceptual modeling of fluid circulation in the area under investigation

# Foundation of geothermometry

Fluid composition and chemical equilibria attained at depth are maintained during transfer to the surface

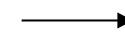
if

*No (or reduced) mixing*

*No (or reduced) dilution*

*No re-equilibration*

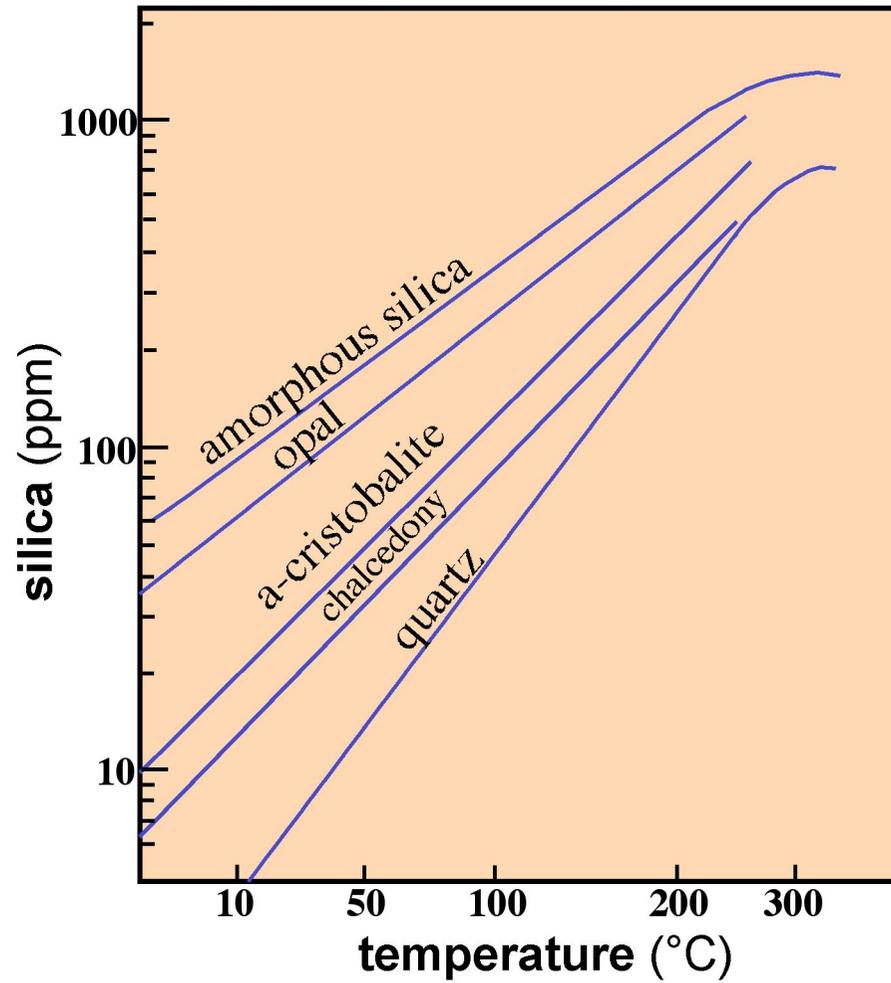
**Conservative elements:** Cl, B, Br,  
(N<sub>2</sub>, He, noble gases)



**tracers**

**Geoindicators (of deep equilibria,  
temperature, pressure):**  
SiO<sub>2</sub>, Na, K., Ca (CO<sub>2</sub>, H<sub>2</sub>, CH<sub>4</sub>)

solubility of silica phases in water at the vapor pressure of solutions with increasing temperature (Fournier, 1991)



$$\text{Quartz-no steam loss } t^{\circ}\text{C} = \frac{1309}{5.19 - \log S} - 273.15 \quad (1)$$

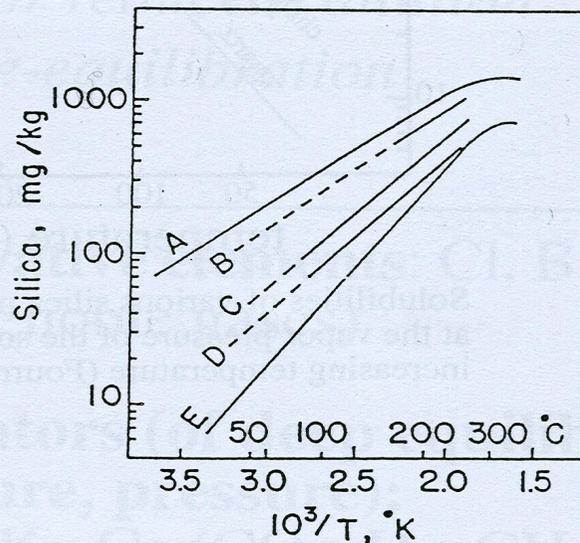
$$\text{Quartz-maximum steam loss at } 100^{\circ}\text{C } t^{\circ}\text{C} = \frac{1522}{5.75 - \log S} - 273.15 \quad (2)$$

$$\text{Chalcedony } t^{\circ}\text{C} = \frac{1032}{4.69 - \log S} - 273.15 \quad (3)$$

$$\alpha\text{-Cristobalite } t^{\circ}\text{C} = \frac{1000}{4.78 - \log S} - 273.15 \quad (4)$$

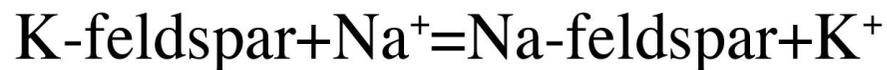
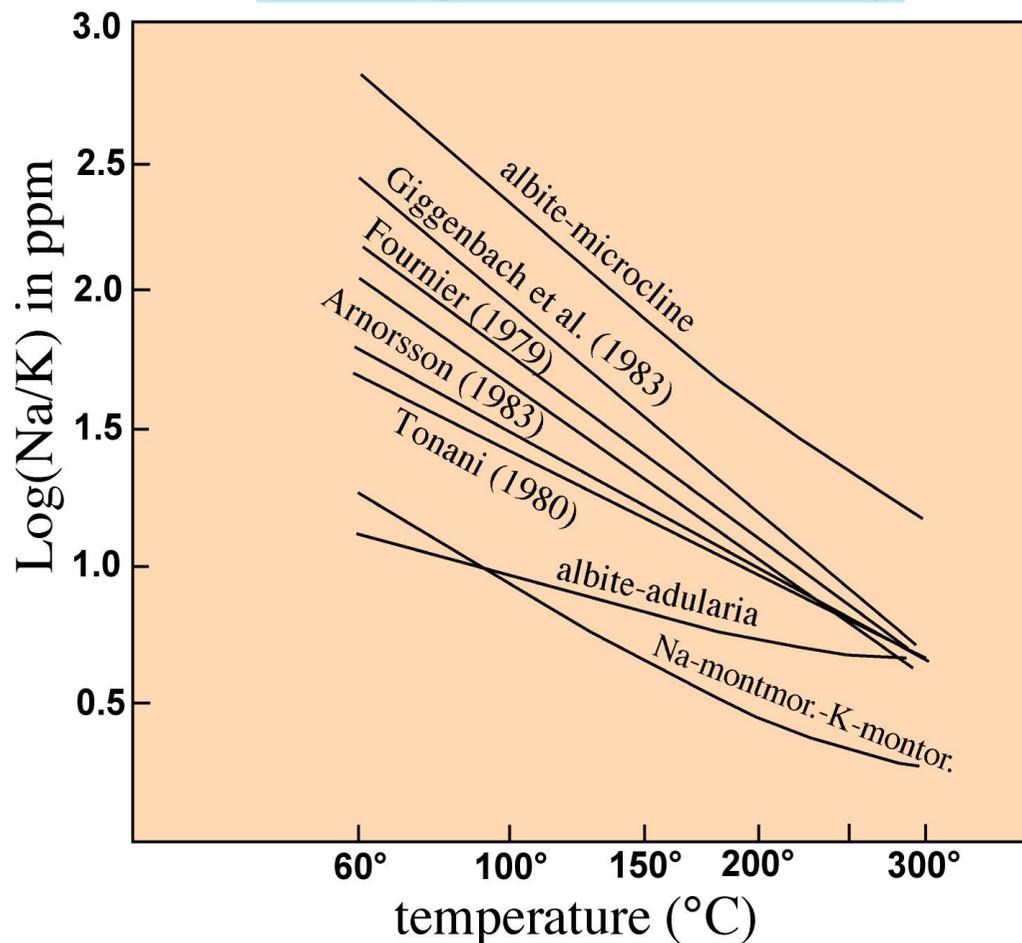
$$\text{Opal-CT } (\beta\text{-Cristobalite}) \quad t^{\circ}\text{C} = \frac{781}{4.51 - \log S} - 273.15 \quad (5)$$

$$\text{Amorphous silica } t^{\circ}\text{C} = \frac{731}{4.52 - \log S} - 273.15 \quad (6)$$



Before applying any relation the n. of silica phases on which the solution is saturated must be calculate (speciation programs; i.e. Wateq, EQ3..etc.); only the still soluble phase (or the more soluble must be used for geothermometry, otherwise you will over estimate deep Temperatures.

## Na/K geothermometry



$$K_{(eq)} = aK^+ / aNa^+$$

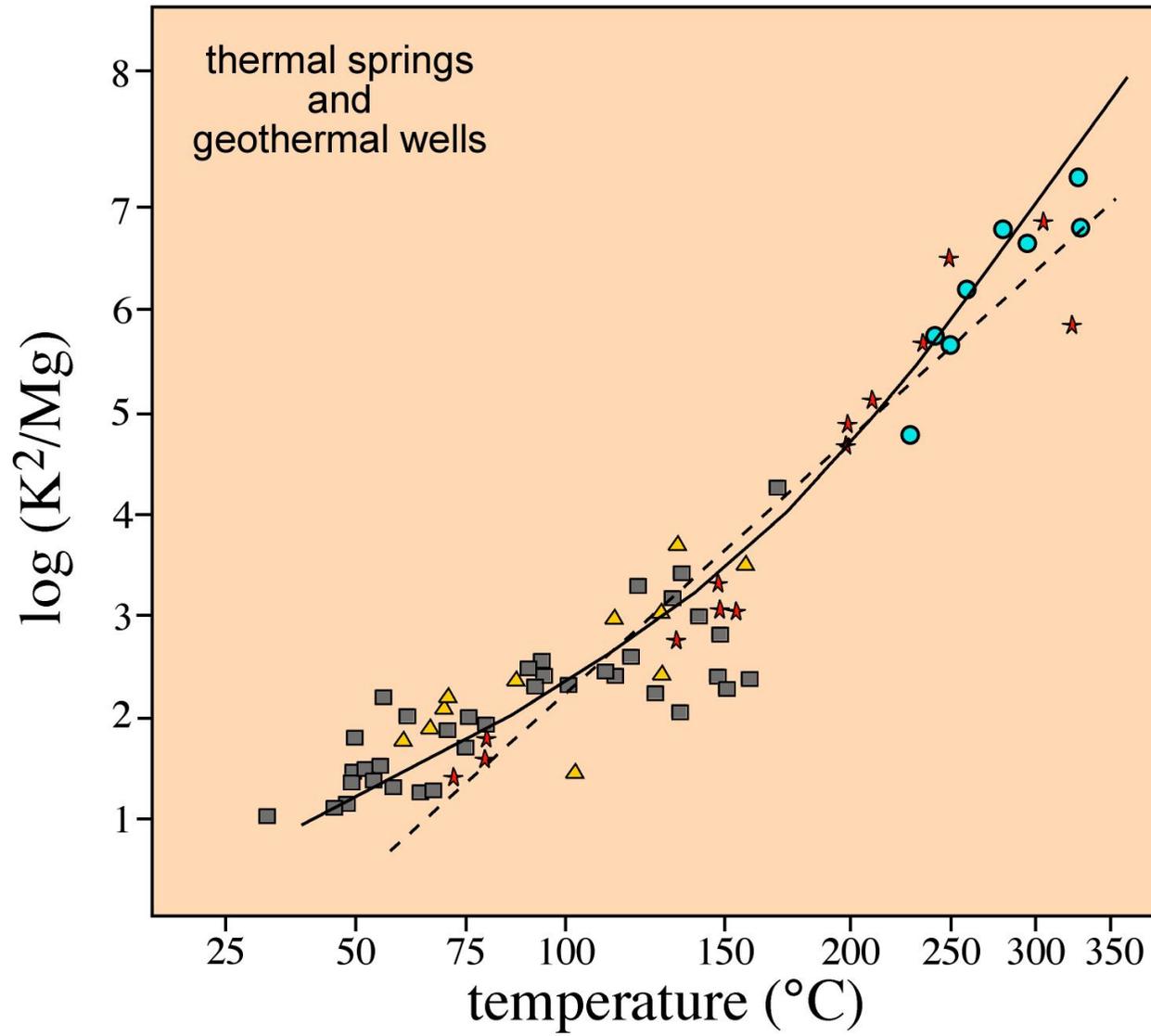
$$\Delta F^\circ = -RT \ln K_{(eq)}$$

$$\Delta F^\circ = \Delta H^\circ - T\Delta S$$

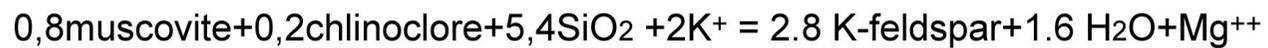
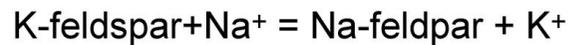
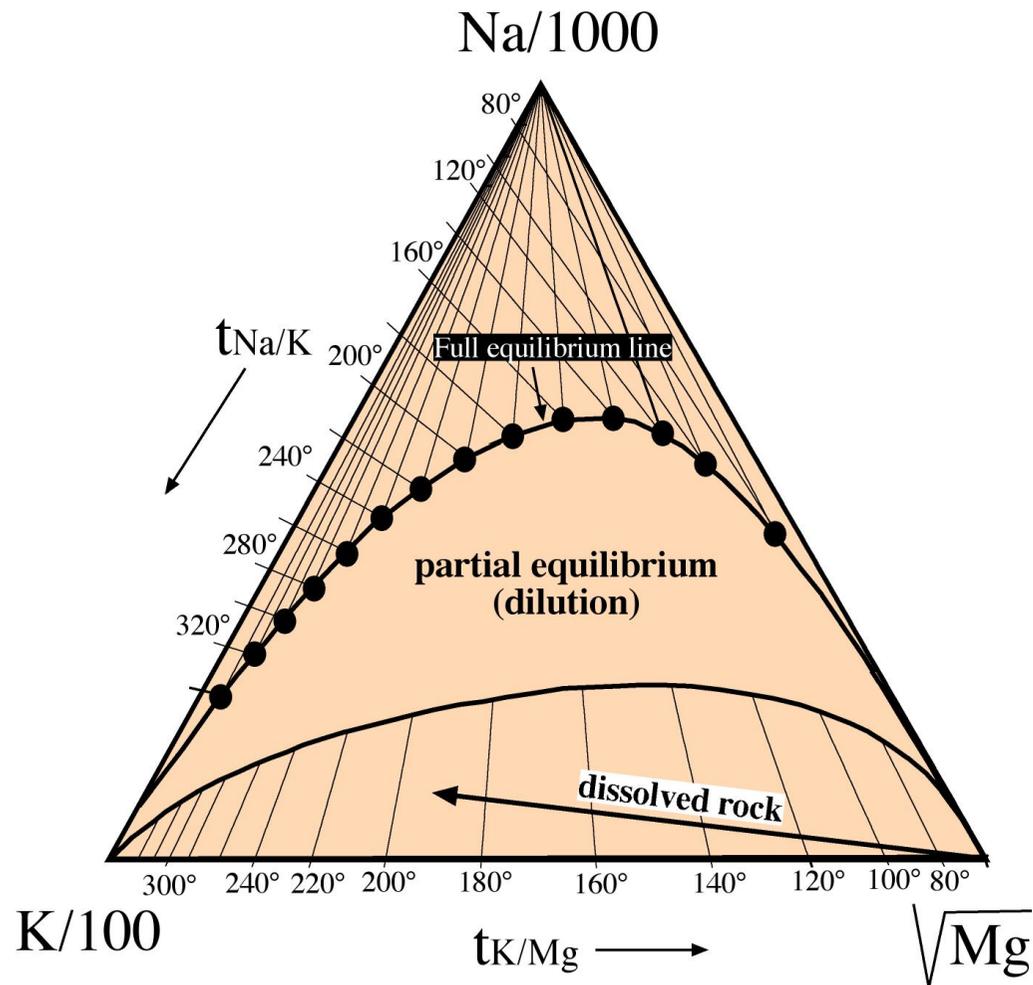
## equations for cation geothermometers (concentrations in mg/kg)

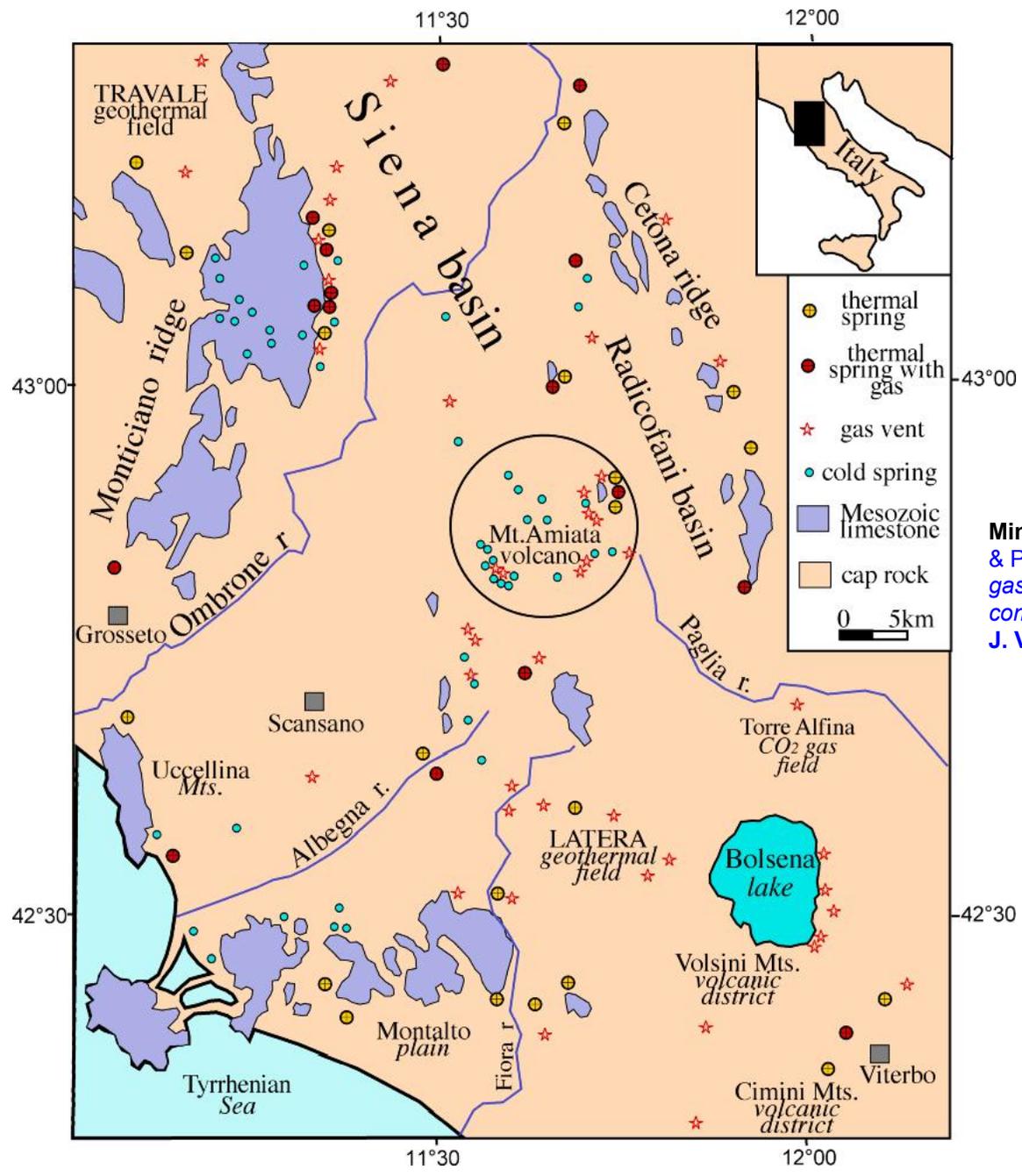
geothermometer	equation	author
Na/K	$t (^{\circ}\text{C}) = \frac{856}{0.857 + \log(\text{Na/K})} - 273.15$	Truesdell (1976)
Na/K	$t (^{\circ}\text{C}) = \frac{833}{0.780 + \log(\text{Na/K})} - 273.15$	Tonani (1980)
Na/K	$t (^{\circ}\text{C}) = \frac{933}{0.993 + \log(\text{Na/K})} - 273.15$	Arnorsson (1983)
Na/K	$t (^{\circ}\text{C}) = \frac{1319}{1.699 + \log(\text{Na/K})} - 273.15$	Arnorsson (1983)
Na/K	$t (^{\circ}\text{C}) = \frac{1217}{1.483 + \log(\text{Na/K})} - 273.15$	Fournier (1983)
Na/K	$t (^{\circ}\text{C}) = \frac{1178}{1.470 + \log(\text{Na/K})} - 273.15$	Nieva & Nieva (1987)
Na/K	$t (^{\circ}\text{C}) = \frac{1390}{1.750 + \log(\text{Na/K})} - 273.15$	Giggenbach (1983)
K-Mg $\log(\text{K}^2/\text{Mg}) > 1.25$	$t (^{\circ}\text{C}) = \frac{2330}{7.350 - \log(\text{K}^2/\text{Mg})} - 273.15$	Fournier (199?)
K-Mg $\log(\text{K}^2/\text{Mg}) < 1.25$	$t (^{\circ}\text{C}) = \frac{1077}{4.033 + \log(\text{K}^2/\text{Mg})} - 273.15$	Fournier (199?)
Li-Mg	$t (^{\circ}\text{C}) = \frac{2200}{5.470 - \log(\text{Li}/\text{Mg}^{1/2})} - 273.15$	Kharaka & Mariner (1989)
Na/Li	$t (^{\circ}\text{C}) = \frac{1590}{0.779 + \log(\text{Na}/\text{Li})} - 273.15$	Kharaka et al. (1982)
Na/Li	$t (^{\circ}\text{C}) = \frac{1000}{0.389 + \log(\text{Na}/\text{Li})} - 273.15$	Fouillac & Michard (1981)
Na/Li	$t (^{\circ}\text{C}) = \frac{1195}{0.130 + \log(\text{Na}/\text{Li})} - 273.15$	Fouillac & Michard (1981)
Na/Ca	$t (^{\circ}\text{C}) = \frac{1097}{3.080 - \log(\text{Na}/\text{Ca}^{1/2})} - 273.15$	Tonani (1980)
K/Ca	$t (^{\circ}\text{C}) = \frac{1930}{3.861 - \log(\text{K}/\text{Ca}^{1/2})} - 273.15$	Tonani (1980)
Na-K-Ca	$t (^{\circ}\text{C}) = \frac{1647}{\log(\text{Na}/\text{K}) + \beta(\log(\text{Ca}^{1/2}/\text{Na}) + 2.06) + 2.47} - 273.15$ $\beta = 4/3$ for $t < 100^{\circ}\text{C}$ $\beta = 1/3$ for $t > 100^{\circ}\text{C}$	Fournier & Truesdell (1973)

# K/Mg geothermometer



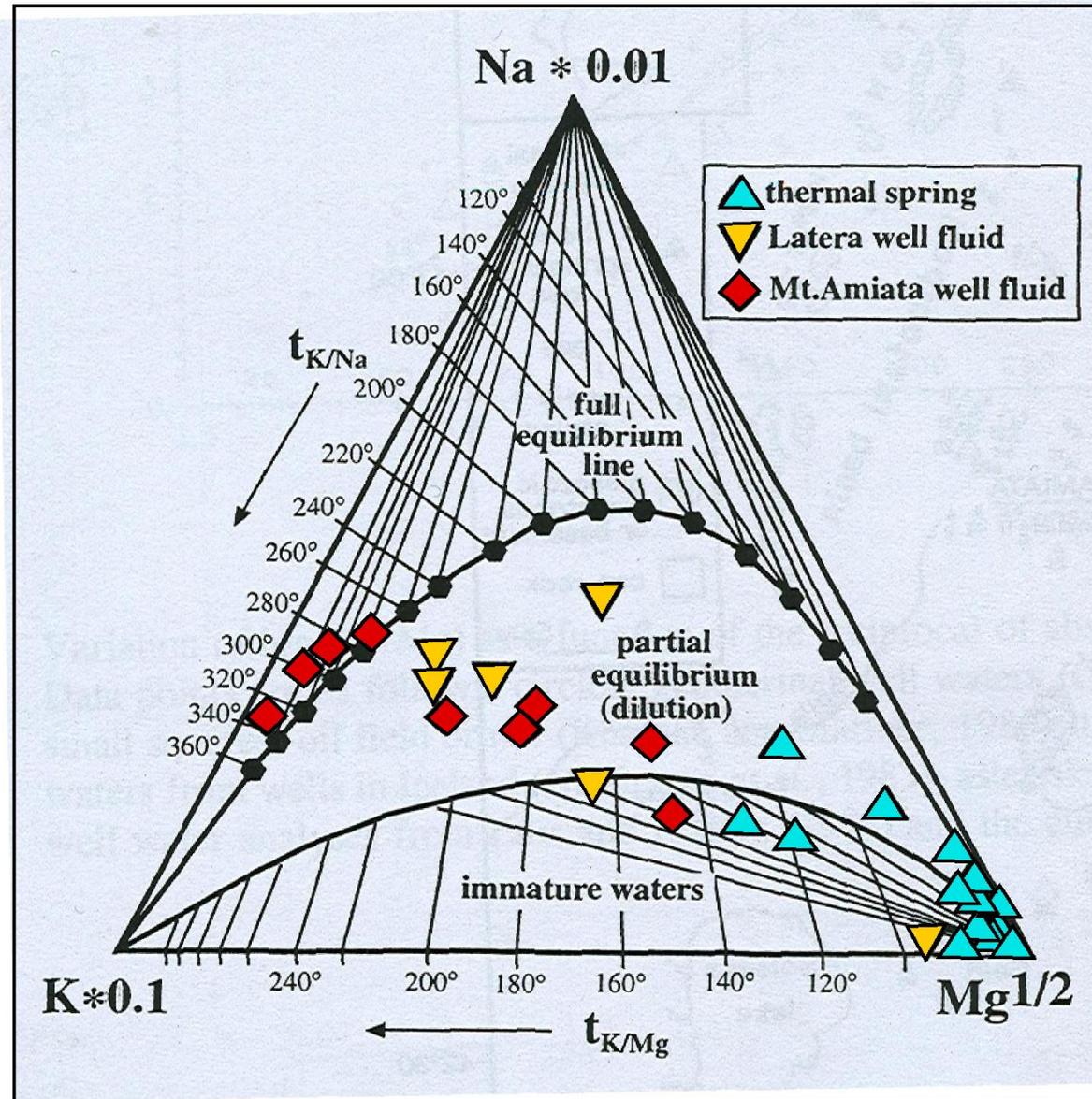
# Evaluation of Na-K-Mg temperatures



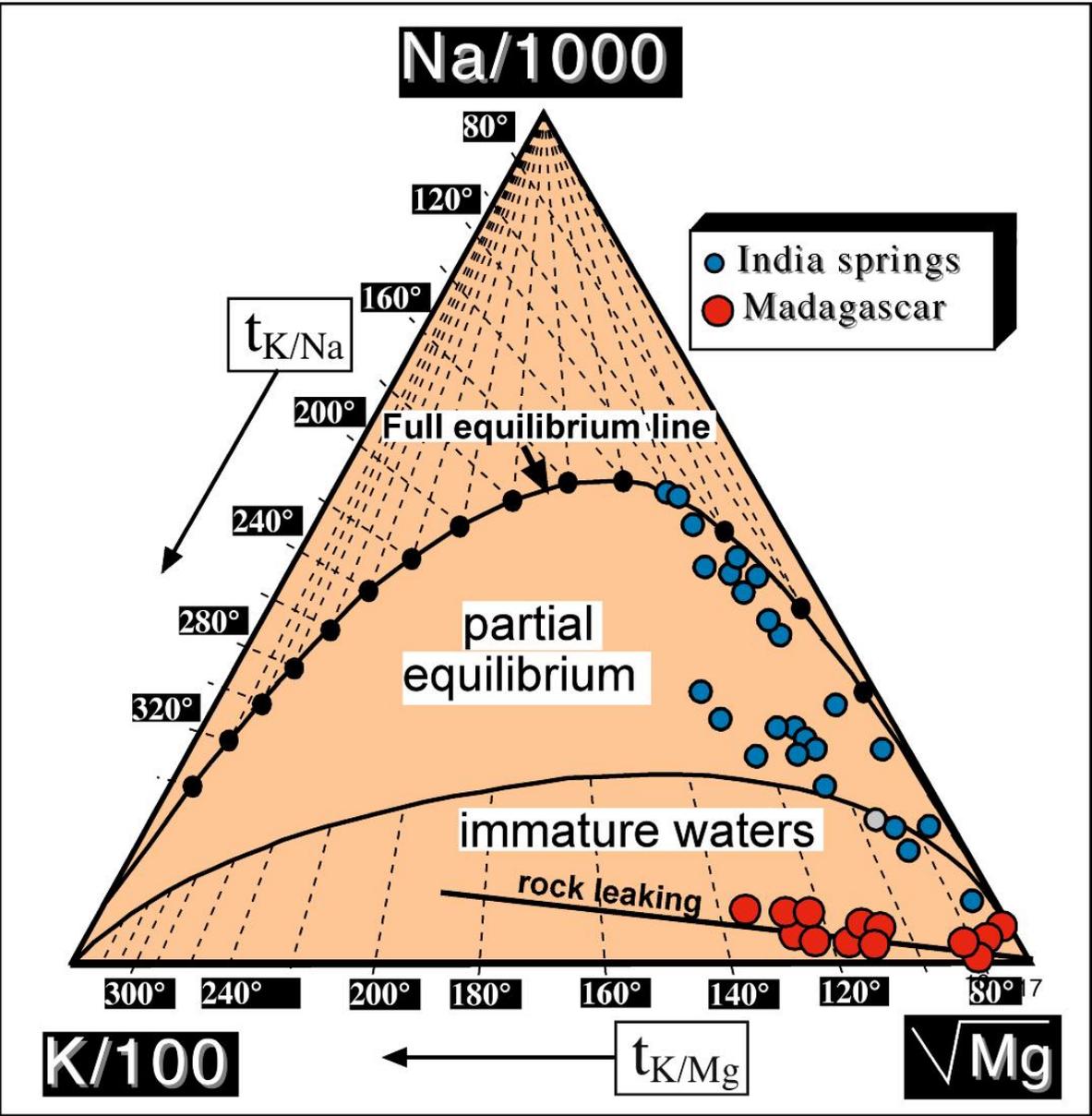


**Minissale A., Magro G., Vaselli O., Verrucchi C. & Perticone I. (1997) Geochemistry of water and gas discharges from the Mt. Amiata silicic complex and surrounding areas (central Italy). *J. Volcanol. Geotherm. Res.* 79, 223-25 1**

# Mt. Amiata geothermal area



# India and Madagascar



Geothermal indicators  
(tracers)

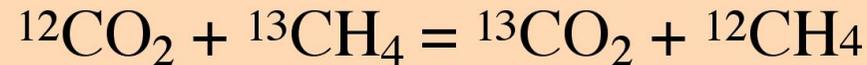
**NH<sub>4</sub> (Hg)**

If steam condenses in shallow aquifer(s), surface (cold) springs should be anomalous in terms of ammonium with respect to the regional background

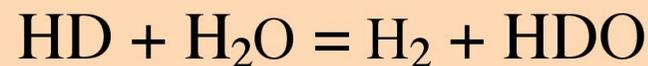
A closed anomaly in shallow aquifers in a active (or Quaternary) volcanic area, possibly associated to a  $p\text{CO}_2$  anomaly, likely represents the outflow path of rising steam (condensation in the aquifer)

*Because isotopes fractionate during chemical or physical processes and that the fractionation factor ( $\alpha$ ) is a function of temperature, several isotopic geothermometers have been proposed*

$$\alpha = \frac{\text{Isotopic ratio in phase or process (1)}}{\text{Isotopic ratio in phase or process (2)}}$$

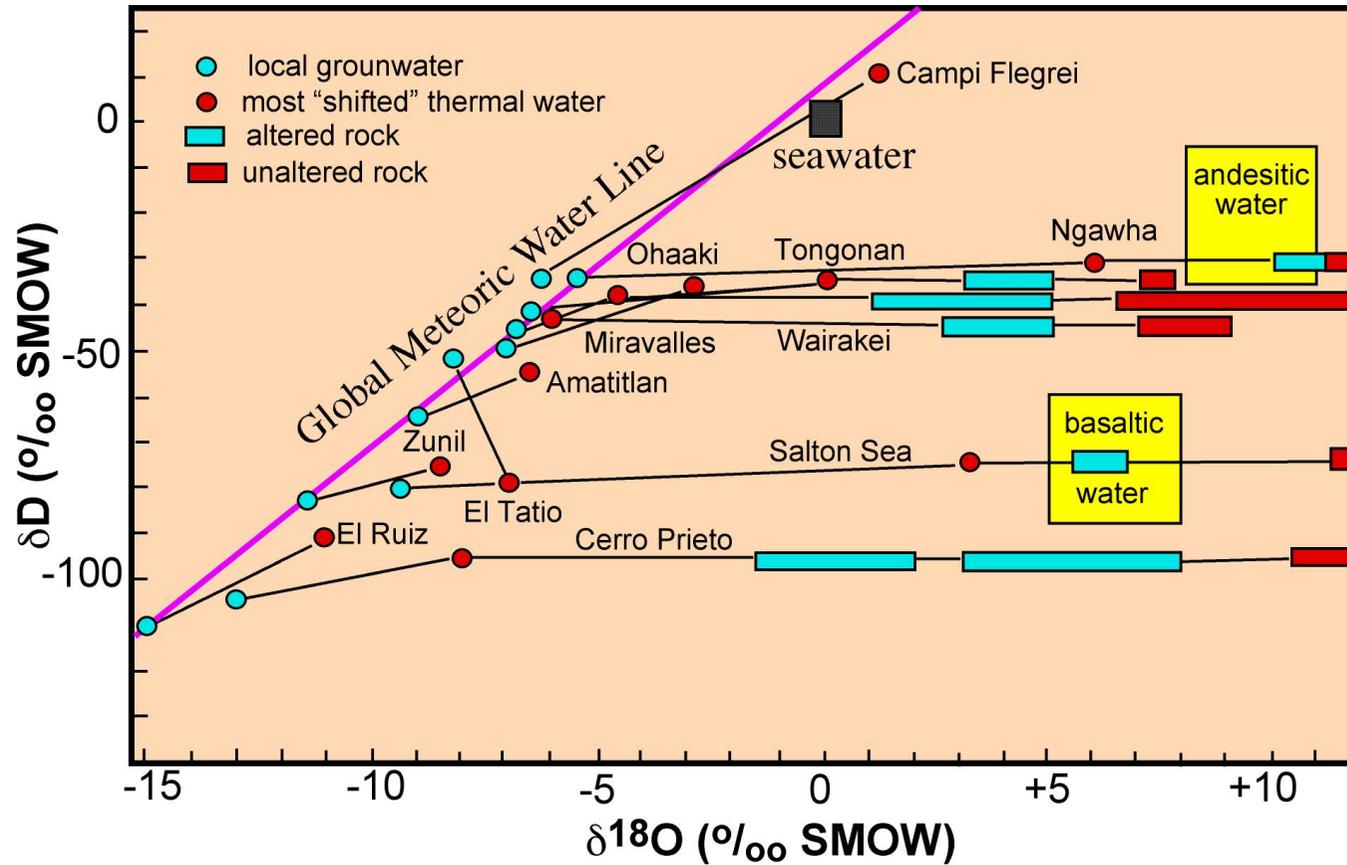


$$1000 \ln\alpha = \frac{15.25 * 10^3}{T(^{\circ}\text{K})} + \frac{2.432 * 10^6}{T(^{\circ}\text{K})} - 9.56$$

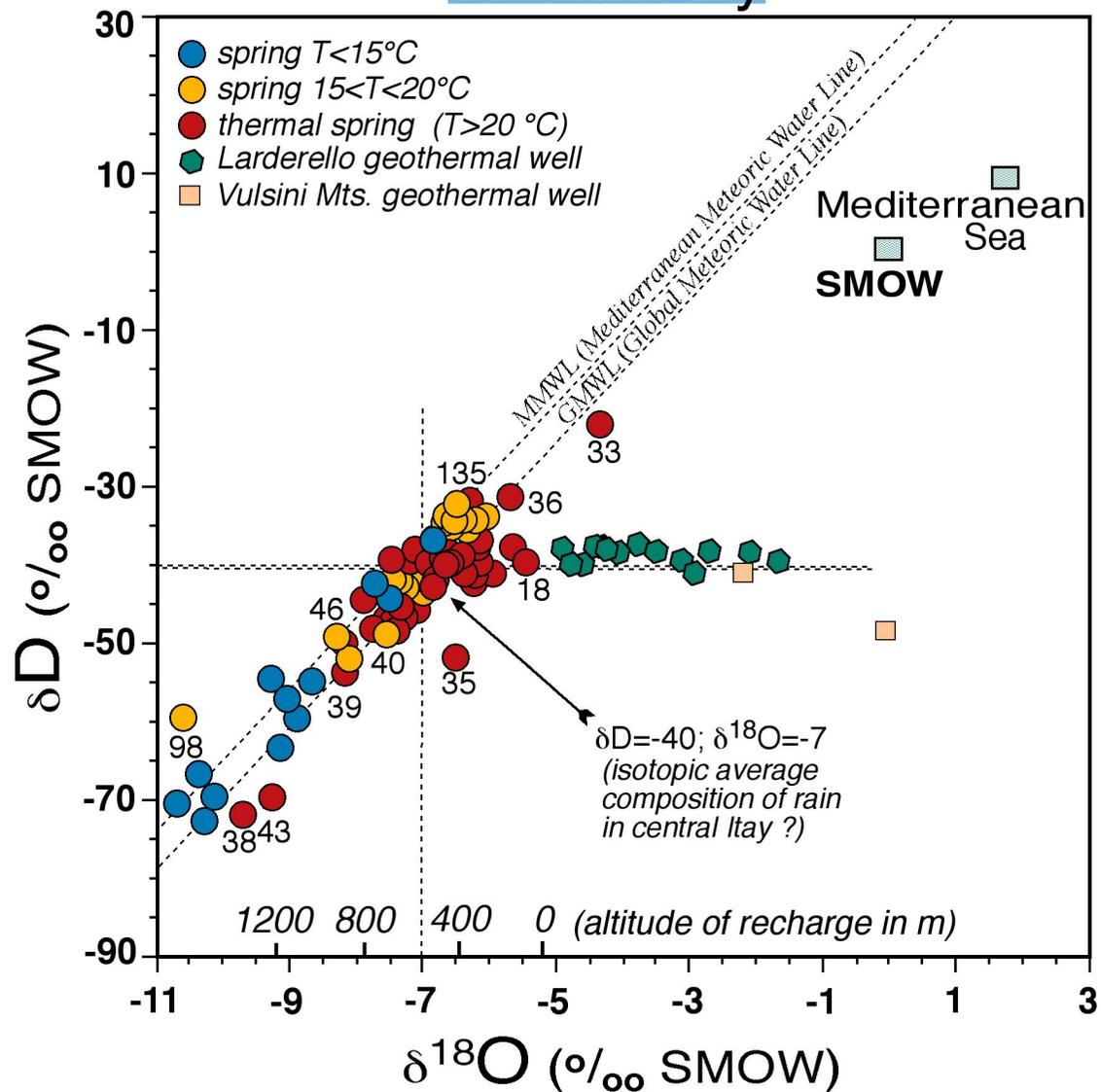


$$1000 \ln\alpha = \frac{396.8 * 10^3}{T(^{\circ}\text{K})} + \frac{11.76 * 10^6}{T(^{\circ}\text{K})} - 217.3$$

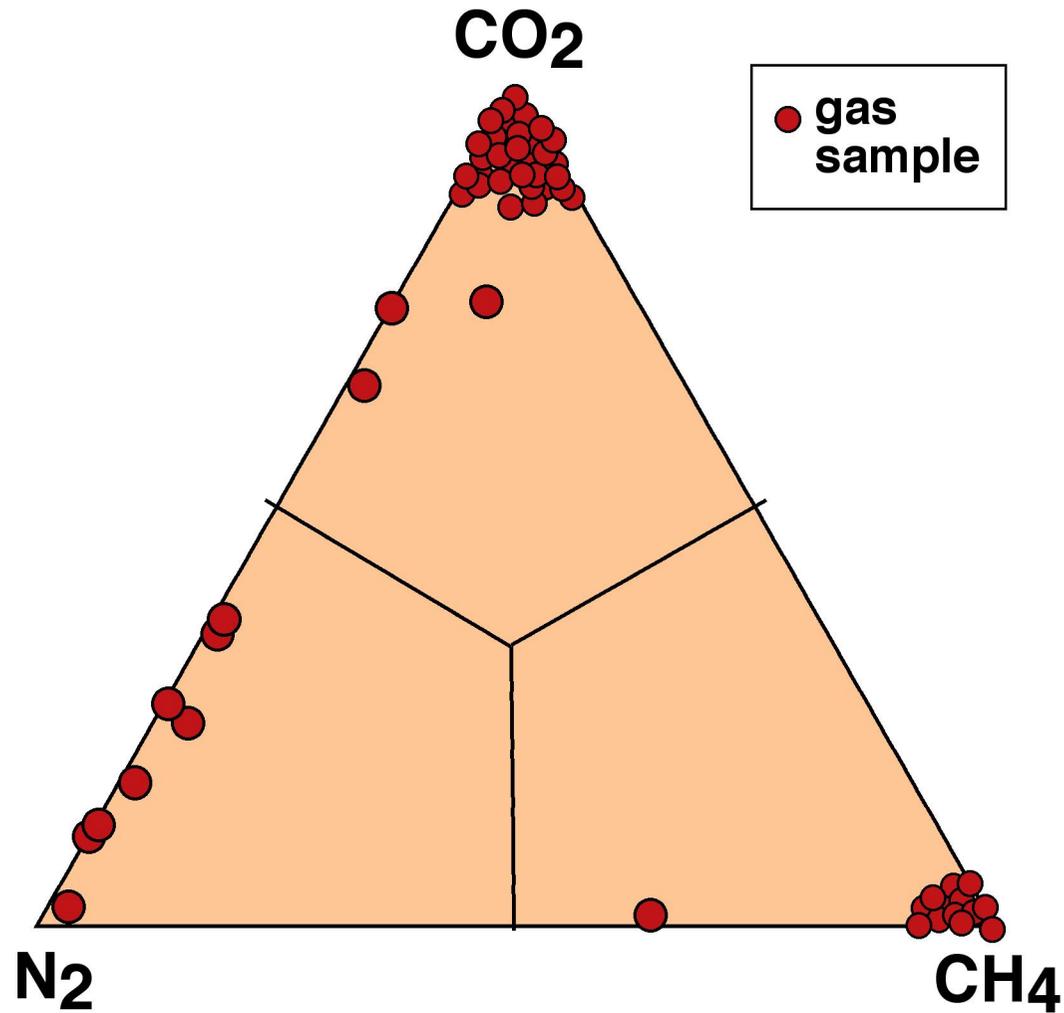
# Oxygen-18 shift with respect to the Global Meteoric Water Line



# central Italy

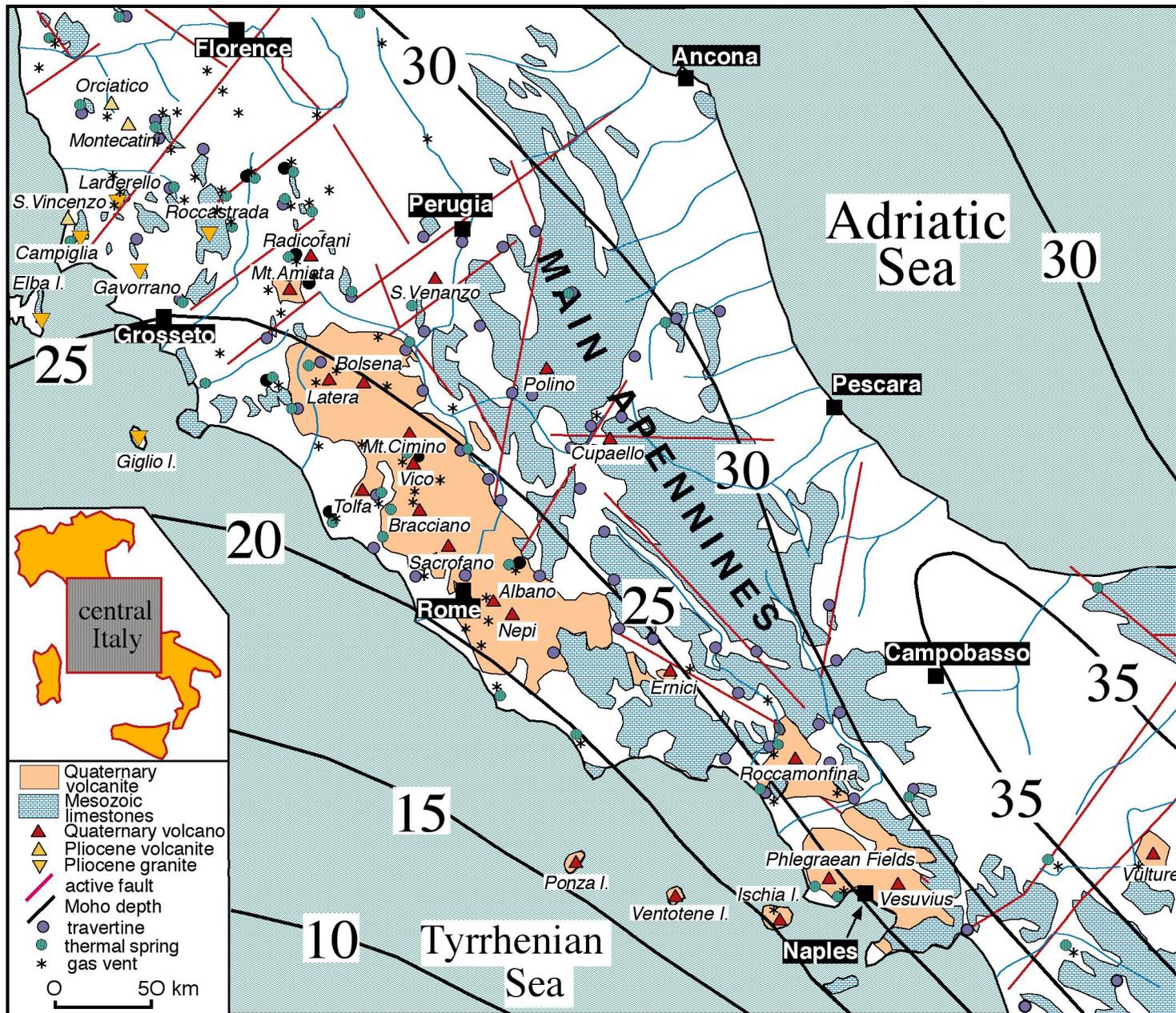


# northern-central Apennines



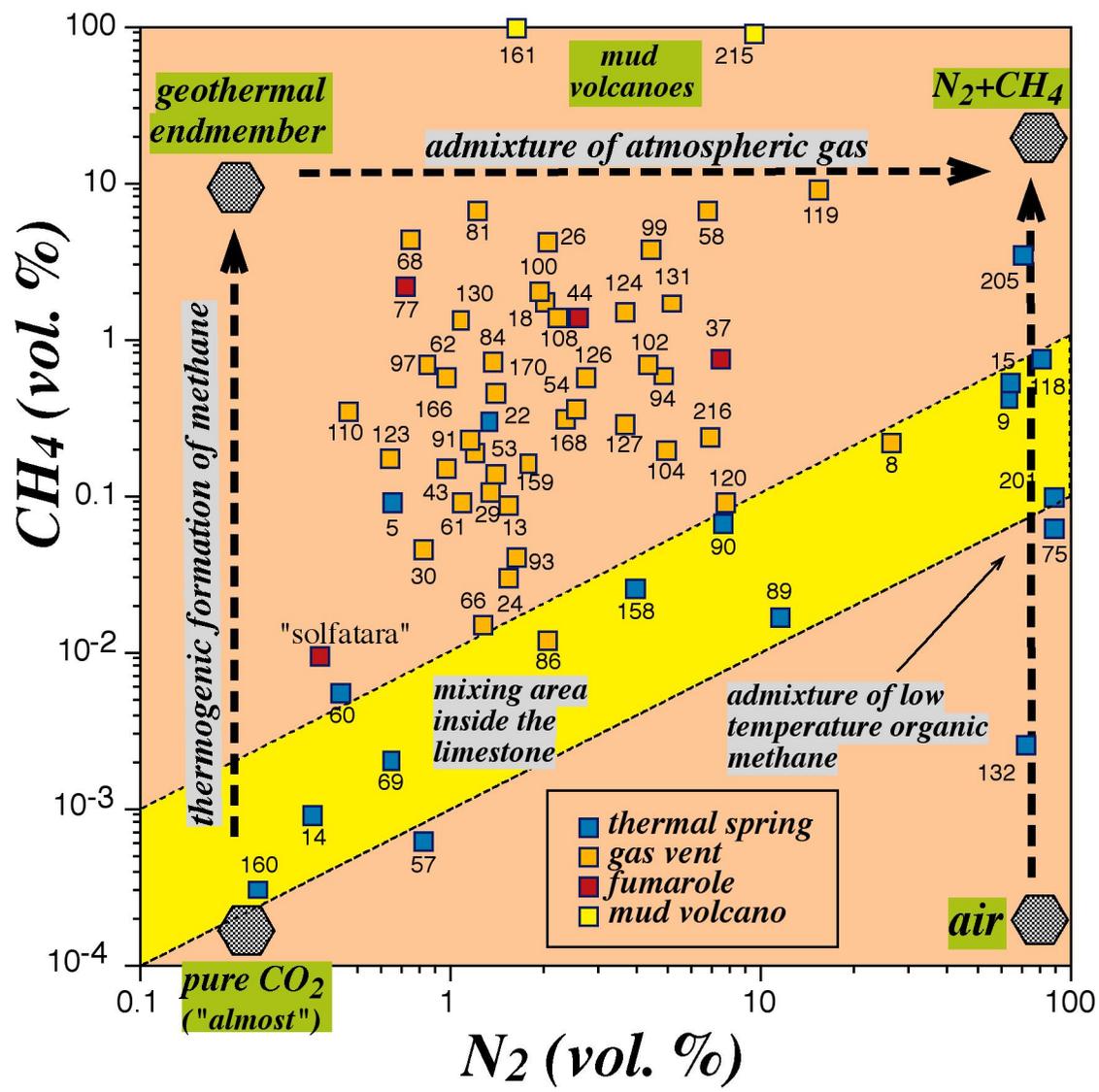
relative abundance of main components in gas vents in central Italy

# schematic map of central Italy







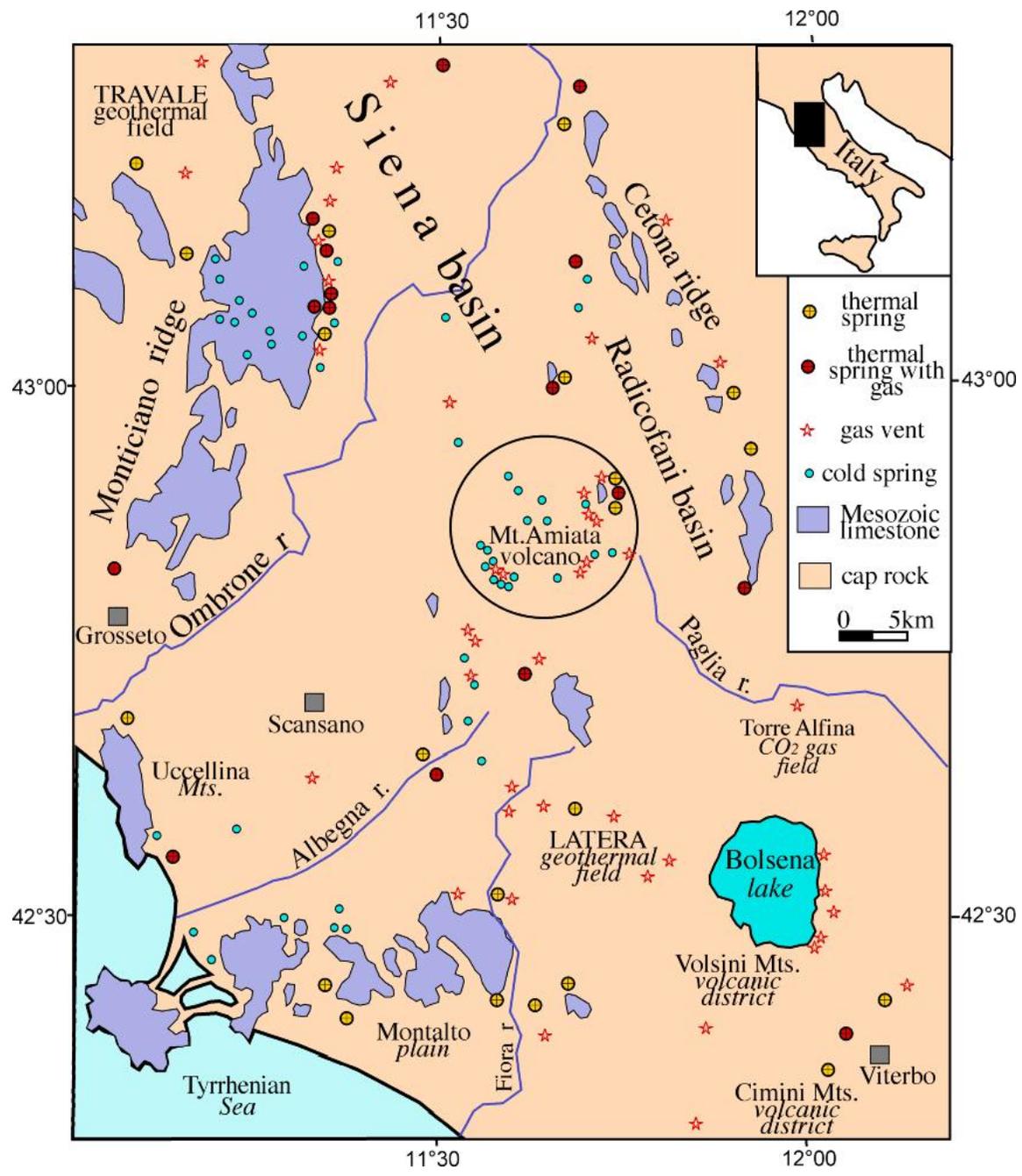


# Gas geothermometers

- 1) Apart from H<sub>2</sub>S and CO<sub>2</sub> partly soluble in water, H<sub>2</sub>, CO, CH<sub>4</sub> and noble gases are poorly soluble, so, they keep recorded their deep equilibrium  $\text{CH}_4 + 2\text{H}_2\text{O} = \text{CO}_2 + 4\text{H}_2$
- 2) Geothermal gases do not undergo dilution with shallow components (apart from CH<sub>4</sub> in Neogene basin).
- 3) Gases do not re-equilibrate at shallow conditions
- 4) Ar and Ne can be used as atmospheric tracers
- 5) Relative ratios in surface gas vents can be used as geothermometers (i.e.:  $\log(X_{\text{CH}_4}/X_{\text{CO}_2}) = 5280/T(^{\circ}\text{K}) - 11.12$ )

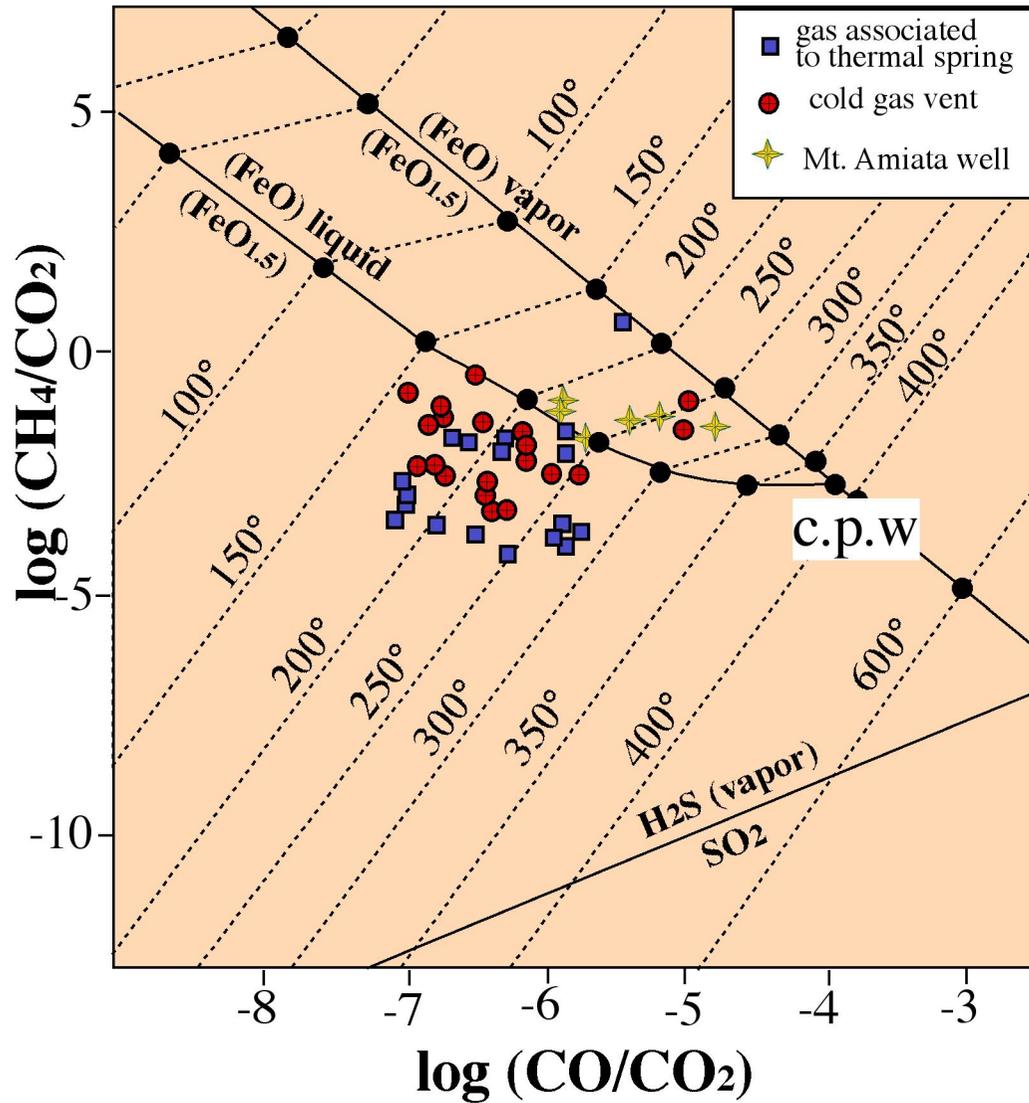
## drawbacks

- 1) Solubilities of gas components in liquid phase are different among them and is variably different at different temperature
- 2) Gas components have different fractionation factors between liquid phase and steam in case of phase separation

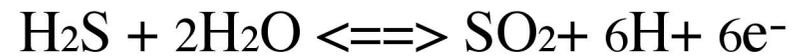
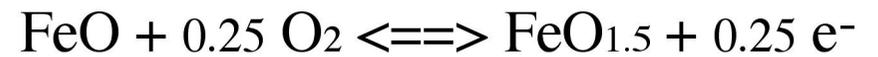
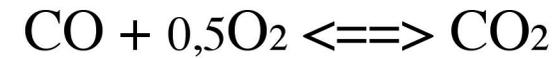
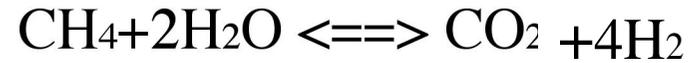


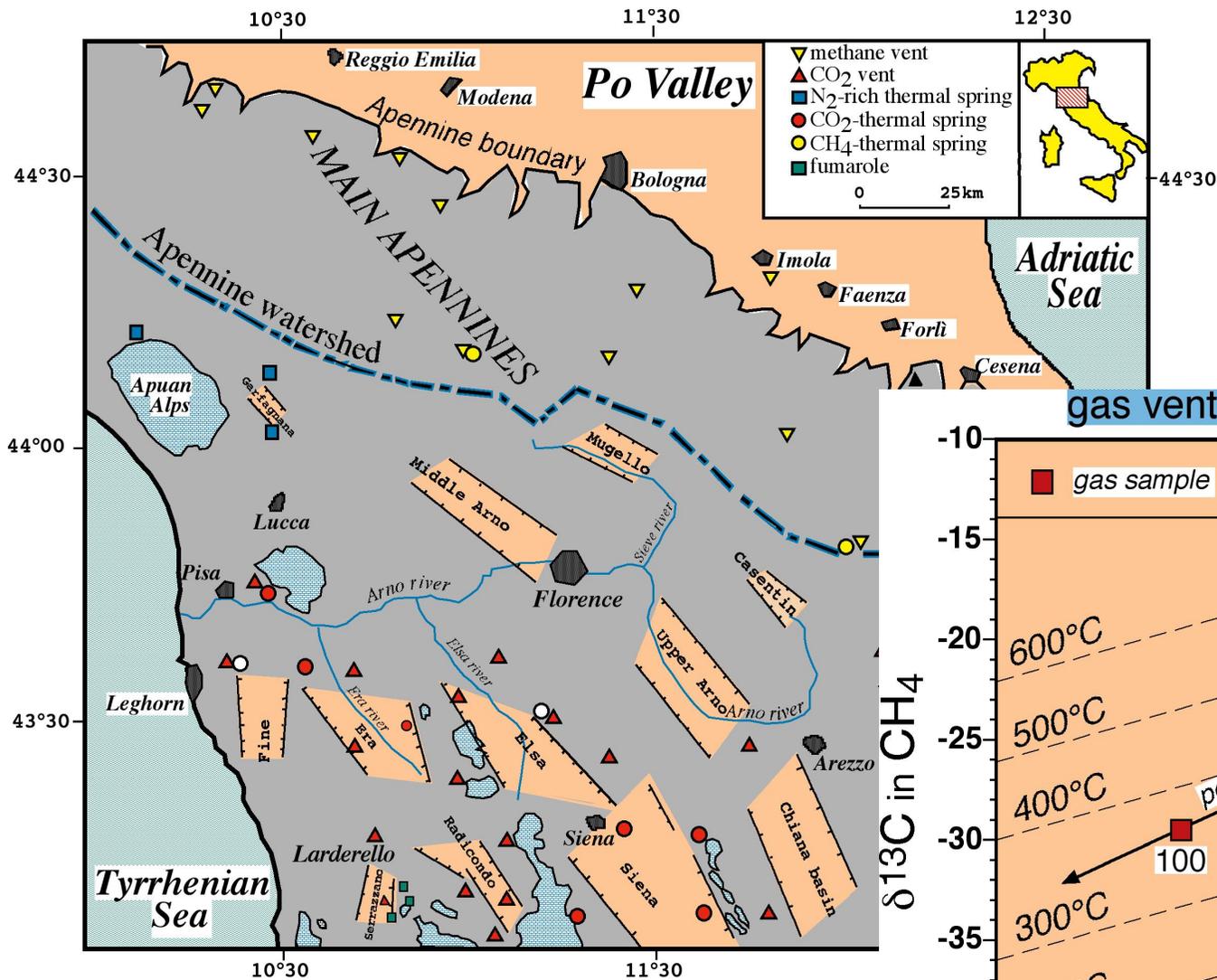
# Mt.Amiata

## geothermometry in the gas phase

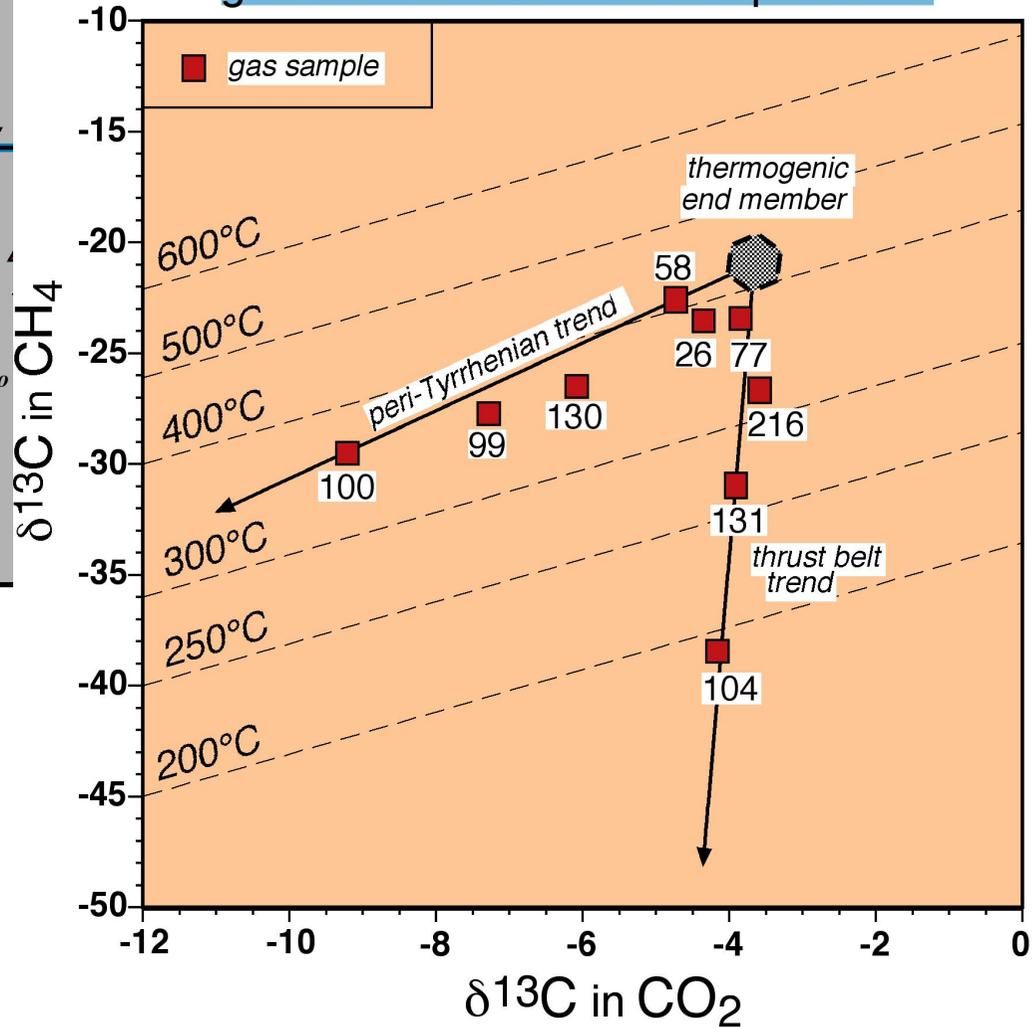


## Mt. Amiata geothermal field





gas vents in the Northern Apennines



# $^3\text{He}/^4\text{He}$

$^3\text{He}/^4\text{He}$  ratio is very sensitive to the presence of mantle magmas residing in the crust

$$R = \frac{^3\text{He}/^4\text{He}(\text{sample})}{^3\text{He}/^4\text{He}(\text{air})}$$

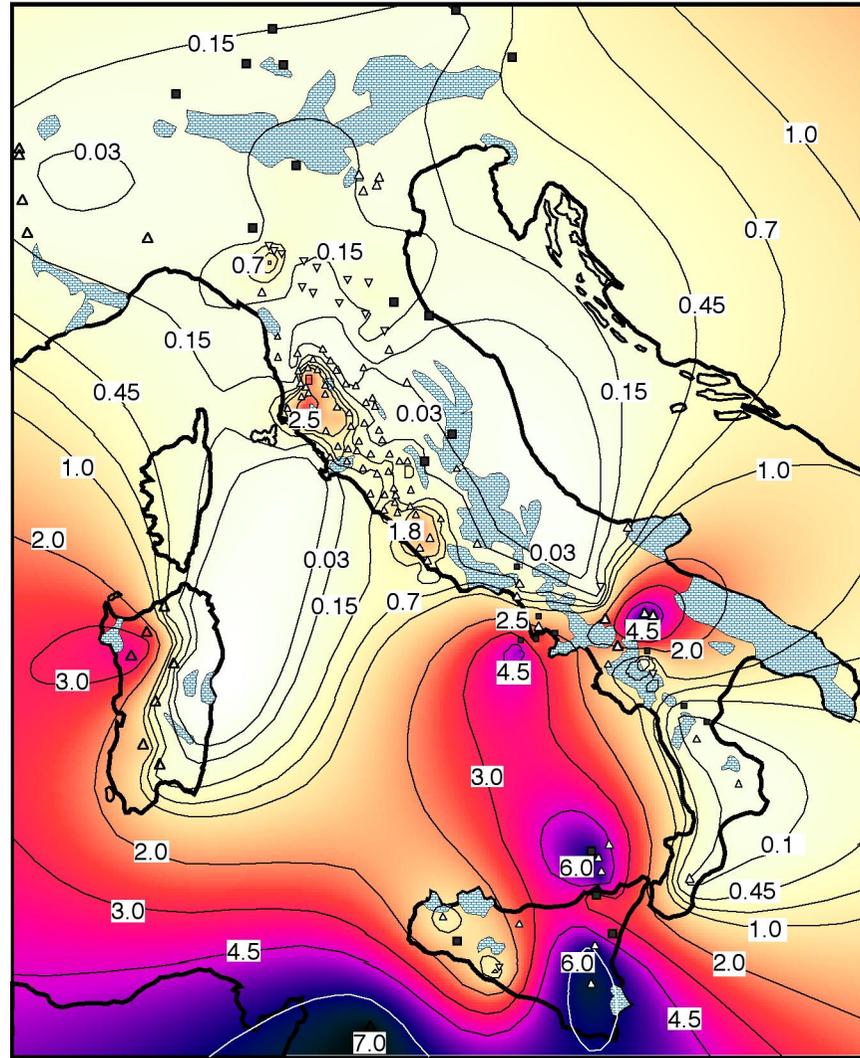
$R \sim 0.005$  in the crust

$R \sim 1$  in the air

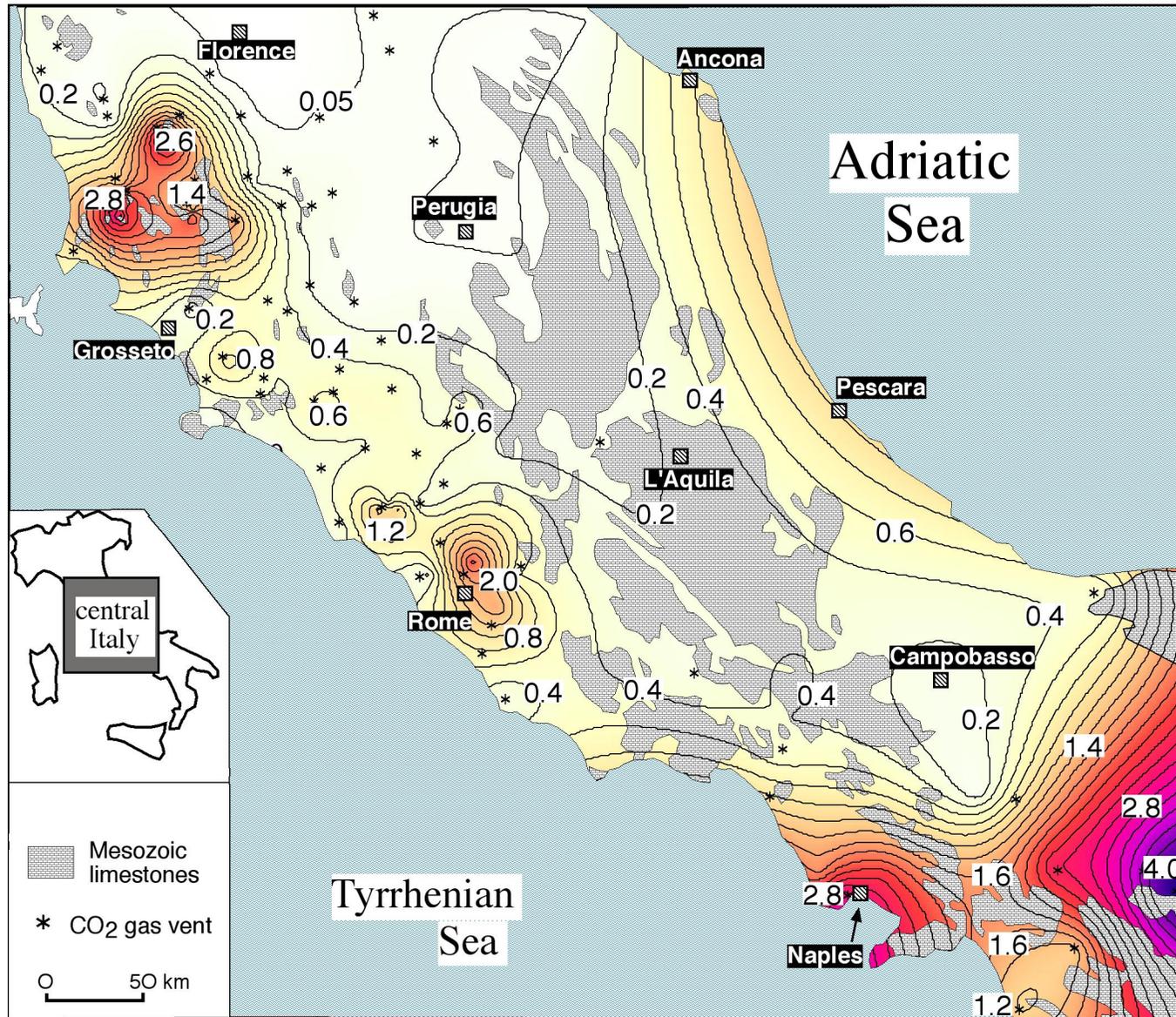
$R \sim 8$  in M.O.R.B.

$R > 15$  in O.I.B.

# helium-3 distribution in gas vents of Italy



# $^3\text{He}/^4\text{He}$ ratio in gas vents in central Italy



## **Electrical conductivity increases:**

- 1) together with temperature
- 2) together with saline fluids  
(geothermal brines are very saline)

The shallower the hot geothermal reservoir the more evident is at the surface the presence of closed conductivity anomalies

## **drowbacks**

Cold saline fluids in Neogene marine and/or continental sediments (unconsolidated)

## Prospection through the heat-flow

$$\Phi_{(c)} = K \, dt/dz$$

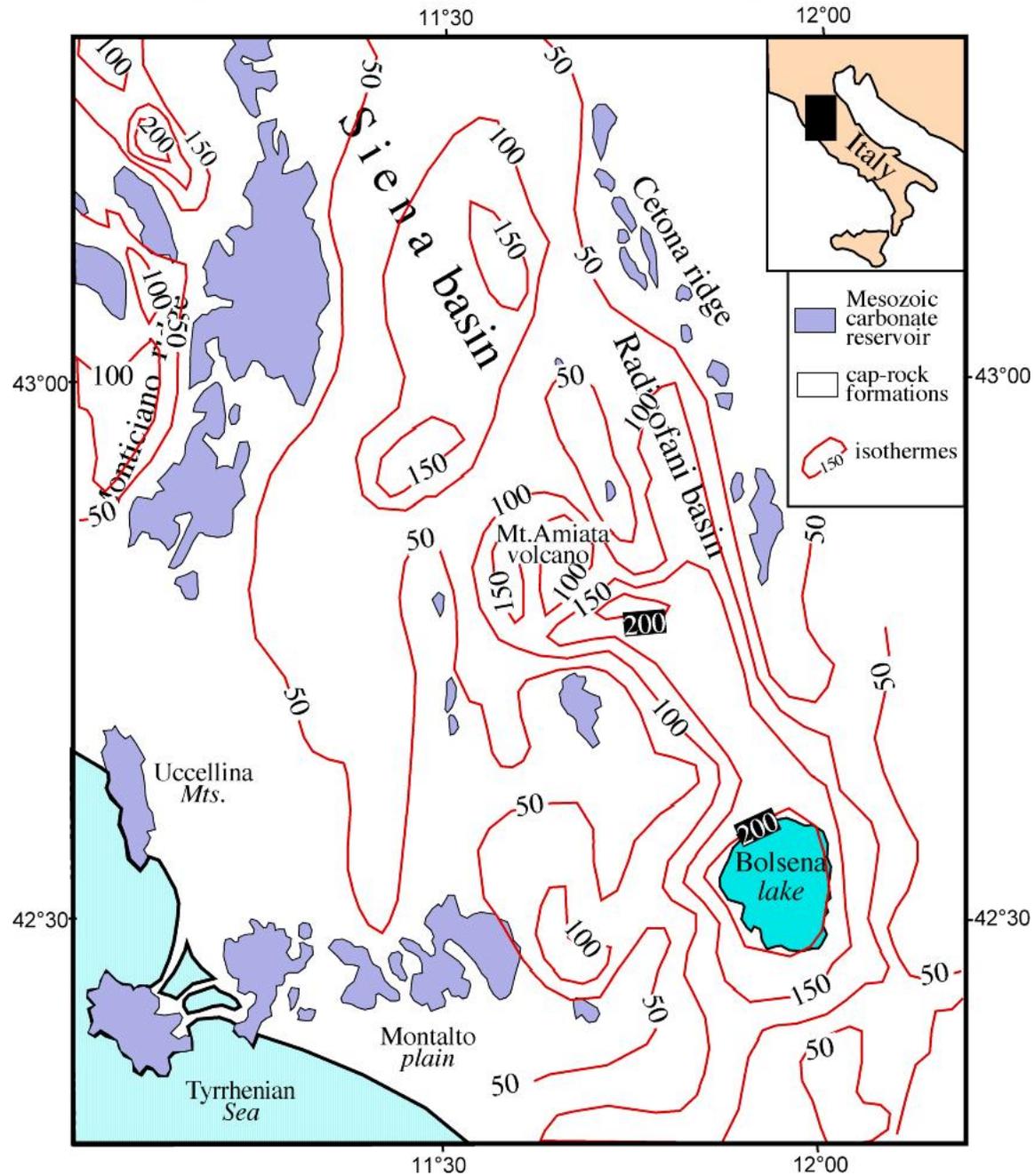
**K= rock thermal conductivity**  
**dt/dz=geothermal gradient**

$\Phi_{(c)} = 60 \text{ mW/m}^2$  in cratonic areas  
 $\Phi_{(c)}$  even  $>400 \text{ mW/m}^2$  in orogenic areas

### drawbacks

- 1)  $\Phi_{(c)}$  in convective systems is near 0
- 2) shallow  $\Phi_{(c)}$  are strongly influenced by the presence of shallow convective relatively thermal aquifers

Estimated temperature at the top of the carbonate formations in the Mt. Amiata region



# Heat-flow in $\text{mW}/\text{m}^2$ in central Italy

